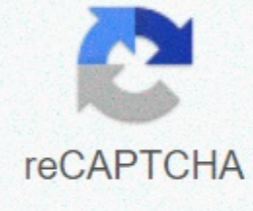




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Writing chemical formulas worksheet answers

Write the correct formula for an ionic compound. Recognize polyatomic ions in chemical formulas. Ionic compounds do not exist as molecules. In the solid state, ionic compounds are in crystal lattice containing many ions each of the cation and anion. An ion formula, such as NaCl , is an empirical formula. This formula simply indicates that sodium chloride is made of an equal number of sodium and chloride ions. Sodium sulfur, another ionic compound, has the formula Na_2S . This formula indicates that this compound is composed of twice as many sodium ions as sulphite ions. This section will teach you how to find the correct ratio of ions, so that you can write a correct formula. If you know the name of a binary ionic compound, you can write its chemical formula. He begins writing the metal ion with his charge, followed by the non-metallic ion with his charge. Since the overall compound must be electrically neutral, decide how many of each ion is necessary for positive and negative expenses to cancel each other. Example 1: Aluminium Nitro and Lithium Oxide Write the formulas for aluminum nitride and lithium oxide. Write the formula for aluminum nitride Write the formula for lithium oxide 1. Write the symbol and charge of the cation (metal) before and the anion (non-metal) second. Al^{3+} : Li^+ : N^{3-} : O^{2-} 2. Use a multiplier to make the total charge of the same cations and anions between them. $1(3+) = 1(3-) + 3 = -3$ total cation charge = total anion charge $2(1+) = 1(2-) + 2 = -2$ 3. Use multipliers as subscribed for each ion. Al_3N_2 : Li_2O 4. Write the final formula. Leave out all expenses and all subscriptions that are 1. AlN : Li_2O An alternative way to write a correct formula for ionic compound is to use the method of cross. In this method, method, the numerical value of each ion charge is crossed to become the subscriber of the other ion. The signs of the charges have dropped. Example 2: The Crisscross method for lead (IV) Oxide Write the formula for lead oxide (IV). Solution Crisscross Method Write the formula for lead oxide (IV) 1. Write the symbol and charge of the cation (metal) before and the anion (non-metal) second. Pb^{4+} : O^{2-} 3. Reduce to the lowest ratio. Pb_2O_4 4. Write the final formula. Leave out all subscribers that are 1. PbO Exercise 2: Write the chemical formula for an ionic compound composed of each pair of ions. the ion of calcium and oxygen ion 2+ copper and sulphurous ion 1+ copper and sulphurous ion Reply to: CaO Reply b: CuS Response c: Cu2S Be aware that ionic compounds are empirical formulas and thus must be written as the lowest ratio of ions. Example 3: Sulfur Compound Write the formula for sodium combined with sulfur. Solution Crisscross Method Write the formula for sodium combined with sulfur 1. Write the symbol and charge of the cation (metal) before and the anion (non-metal) second. Na^+ : S^{2-} 3. Reduce to the lowest ratio. This step is not necessary. 4. Write the final formula. Leave out all subscribers that are 1. Na_2S Exercise 3: Write the formula for each ionic compound. Magnesium Hydrochloride Sodium Bromide Response to: NaBr Response b: LiCl Response MgO Some ions consist of groups of atoms tied together and have a total electrical charge. Why these ions ion more than an atom, are called polyatomic ions. polyatomic ions have formulas, names and features charges that must be stored. for example, NO_3^- is the nitrate ion; has a nitrogen atom and three oxygen atoms and a total charge of 1-. The table 1 lists the most common polyatomic ions. table 1: polyatomic ions ammonium formula NH_4^+ acetate ion $\text{C}_2\text{H}_3\text{O}_2^-$ (also written CH_3CO_2^-) carbonate ion CO_3^{2-} chromate ion CrO_4^{2-} dichromate ion $\text{Cr}_2\text{O}_7^{2-}$ hydrogen carbonate ion (bicarbonate ion) HCO_3^- the rule for the construction of formulas for ionic compounds containing polyatomic monons is the same as formulas if more than one of a particular polyatomic ion is necessary to balance the charge, the whole formula for the polyatomic ion must be enclosed in parentheses, and the numerical undersigned is placed outside the brackets. This means that the undersigned applies to the entire polyatomic ion. an example is $\text{Ba}(\text{NO}_3)_2$. writing a formula for ionic compounds containing polyatomic ions also involves the same passages as a binary ionic compound. write the symbol and charge of the cation followed by the symbol and the charge of the anion. example 4: calcium nitrates write the formula for calcium nitrate. crisscross method solution write formula for calcium nitrate 1. write the symbol and charge of the cation (metal) before and the anion (non-metal) second. Ca^{2+} : N^{3-} 2. transpose only the number of positive charge to become the subscriber of the anion and the only number of the negative charge to become the of the cation. 3. Reduce to the lowest ratio. 4. Write the final formula. Leave out all subscribers that are 1. If there is only 1 of the polyatomic ion, leave the brackets. $\text{Ca}(\text{NO}_3)_2$ Example 5: Write the chemical formula for an ionic compound composed of potassium ions and sulphate ions. Potassium ions have a charge of 1+, while sulphate ions have a charge of 2-. We will need two potassium ions to balance the charge on the sulphate ion, so the correct chemical formula is K_2SO_4 . Exercise 5: Write the chemical formula for an ionic compound composed of each pair of ions. Magnesium Ion and Carbonated Ion Aluminum Ion and Acetate Ion Response to: $\text{Al}(\text{CH}_3\text{COO})_3$ There are two ways to recognize ionic compounds. First, metal compounds and non-metallic elements are usually hyonic. For example, CaBr_2 contains a metal element (calcium, a group 2 [or 2A] metal) and a non-metallic element (bromine, a group 17 [or 7A] non-metallic). Therefore, it is most likely an ionic compound. (It's actually ionic.) On the contrary, the NO_2 compound contains two elements that are both non-metallic (nitrogen, group 15 [or 5A], and oxygen, from group 16 [or 6A]). It is not an ionic compound; it belongs to the category of covalent compounds discussed elsewhere. Also note that this combination of nitrogen and oxygen has no specified electrical charge, so it is not nitrite ion. Secondly, if you recognize the formula of a polyatomic ion in a compound, the compound is ionic. For example, if you see the $\text{Ba}(\text{NO}_3)_2$ formula, you can recognize the "NO3" part as a nitrate ion, NO_3^- . (Remember that the convention for writing formulas for ionic compounds is not to include ion charge.) This is a clue that the other part of the formula, Ba, is actually the Ba^{2+} ion, with the charge 2+ balancing the overall charge 2- from the tonitrates. So, this compound composed of ionic too. Example 6: Identify each compound as ionic or non-ionic. Solution Explanation Answer a. Sodium is a metal, and oxygen is a non-metallic. Therefore, the Na_2O should be ionic. Na_2O , b ionic both phosphorus and chlorine are nonmetal. Therefore, PCl_3 is not ionic. PCl_3 , not ionic c. The NH_4 in the formula represents ammonium ion, NH_4^+ , which indicates that this compound is ionic. NH_4Cl , ionic d. Both oxygen and fluorine are not metal. Therefore, OF_2 is not ionic. OF_2 , Ionic exercise 6: Identifies each compound as ionic or non-ionic. N_2O FeCl_3 $(\text{NH}_4)_3\text{PO}_4$ SOCl_2 Reply to: non-ionic Reply b: ionic Reply c: ionic Reply d: non-ionic The formulas for ionic compounds contain symbols and the number of each atom present in a compound in the lowest integer ratio. 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